Hydrogen potential, potential of the calomel electrode and the resulting measuring potential for temperatures between 15°C and 30°C in buffer pH 2

The figures for the H_2-potential are calculated with the Nernst equation. The figures of the calomel electrode are taken from
D. Dobos, Electrochemical Data, 1975, Elsevier.

Temp. °C	H_2-Potential @ pH 2 in Volt	Hg/Hg_2Cl_2 sat. KCl in Volt	Hg/Hg_2Cl_2 vs. H_2 Potential in Volt
15	-0.1144	0.2503	0.3647
16	-0.1148	0.2497	0.3645
17	-0.1152	0.2490	0.3642
18	-0.1156	0.2483	0.3639
19	-0.1160	0.2477	0.3637
20	-0.1163	0.2471	0.3634
21	-0.1167	0.2464	0.3631
22	-0.1171	0.2458	0.3629
23	-0.1175	0.2451	0.3626
24	-0.1179	0.2445	0.3624
25	-0.1183	0.2438	0.3621
26	-0.1187	0.2431	0.3618
27	-0.1191	0.2425	0.3616
28	-0.1195	0.2418	0.3613
29	-0.1199	0.2412	0.3611
30	-0.1203	0.2405	0.3608

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